soda, which would remain in solution in the sea. But the carbonic acid of the atmosphere, having a more powerful affinity for these alkalies than the silica, would wrest them from combination with the silica, as already stated, and would form carbonates of potash (CO^2KO) and soda (CO^2NaO), while the silica would be added to the quartzose rocks of the globe. These carbonates, whether formed in the ocean or on the hill-sides, would, when transported to the ocean, find themselves confronted with chlorid of calcium (ClCa), and probably other chlorids. Chlorid of calcium, carbonate of potash (CO^2KO), and carbonate of soda (CO^2NaO), brought face to face, would immediately enter into arrangements for an exchange of partners. Carbonic acid (CO^2) would incontinently abandon potash (KO) and soda (NaO), and betake itself to calcium (Ca), changing its name, by the aid of a little oxygen, to "lime" (CaO), and forming a union known as carbonate of lime (CO^2CaO). With equal celerity, chlorine (Cl), dispossessed of its calcium (Ca), would compensate itself by seizing upon potash (KO) and soda (NaO), and, after eliminating the oxygen (O) in their constitution, would unite with potassium and sodium, forming chlorid of potassium (ClK) and chlorid of sodium (ClNa). Thus all parties would be better satisfied, and each would abide in its appropriate place. Carbonate of lime (CO^2CaO) refusing, for the greater part, to be dissolved in sea-water, would settle to the bottom and become limestone; while chlorid of sodium (ClNa)—which is only the chemist's name for "common salt"—remained in solution, and thus gave its characteristic salinity to the sea. Chlorid of potassium (ClK) also continues to exist in sea-water in smaller quantity.

The diagram on the following page is intended to represent to the eye the chemical reactions above described. The symbols are familiar to the chemical reader; but they