a little ingenuity, it would, perhaps, not be difficult to represent mechanically.*

Precisely the same laws of union, may be supposed to prevail among the molecules of bodies themselves, as they actually exist around us. Thus, let us take the crystal of oxalic acid, as an instance for illustration. This acid is composed, according to the present language of chemists, of two molecules of carbon, and three of oxygen, which by combining, form the acid; while, to complete the compound molecule, and to adapt it for crystallization, three molecules of water are required to be somehow associated, with each of the molecules of the acid. Now in this case, we suppose, that the two molecules of carbon, (each of which is perhaps already made up of several sub-molecules,) are associated together into one symmetrical super-molecule; that the three molecules of oxygen, associated in a similar manner, are then combined *chemically* with the super-molecule of carbon, and thus form by their union a molecule of oxalic acid; finally, that the three molecules of water are associated into one super-molecule, which unites

* When bodies, as, for example, water, are subjected to intense degrees of heat, it is not improbable that in many instances the self-repulsive molecules are, more or less, separated into their constituent sub-molecules; in which case, of course, the bodies may be supposed to exhibit altogether different elastic powers, and laws of expansion.