

rature, contain an equal number of self-repulsive molecules, the molecules of different gaseous bodies must also have different weights; which weights will be as the specific gravities of the gases; and may be represented by numbers proportional to these specific gravities.

3. From the above relation between the volumes and the weights of bodies in the gaseous state, it follows, that *all bodies must combine with reference to their weights*; that is to say, that the same weight of the same body, (or half or twice as much, &c.) must always combine with the same weight, (or half or twice as much, &c.) not only of the same, but of every other body.

4. The numbers representing the relations among the specific gravities of bodies in the gaseous state, are called the *molecular, or atomic weights*, of the different bodies.

Such is the foundation of what is usually called the *Atomic Theory*; the principles of which theory, are generally acknowledged to regulate chemical combinations.

We shall now conclude the present Treatise on chemistry, with a few remarks, more especially relating to the object of these volumes. And here it may be observed, once for all, that throughout the preceding pages, as well as in what follows, we have endeavoured to state each argument as distinctly as possible, without encumbering it too much with details—in short, to