is introduced into different bodies, they exhibit very different *apparent* temperatures. This property, as possessed by different bodies, constitutes what is called by chemists, their *capacity for heat*; or, their *specific heat*. Thus, if the same quantity of heat, which we supposed to have been introduced into two molecules of ice, had been introduced into two molecules of silver; the apparent temperature of the molecules of silver would have been raised upwards of ten times more, than the apparent temperature of the molecules of ice; hence, in the case of the ice, some of the heat must have disappeared, or become, in the language of chemists, *latent*. The latency of heat appears to depend upon two different causes; or rather, properly speaking, there are two distinct modes of the latency of heat; as may be thus illustrated. Let us take the two bodies above mentioned—ice and silver: these bodies, under the same volume, contain very unequal portions of matter; the silver being ten times as heavy as the ice. The vacuities in the ice therefore, must be very much greater than the vacuities in the silver; hence, when equal quantities of any matter capable of occupying such vacuities, as heat may be supposed to be, are introduced equally into both these bodies, very dissimilar apparent effects must ensue. The more porous body will absorb, and condense within its vacuities, the matter